

## Standardization of NaOH Pre-Lab

1. What does the equivalence point mean? How is the equivalence point found?
2. What is the proper procedure for filling a buret?
3. What reaction will take place to standardize NaOH?
4. How does the indicator let you know that a titration has finished?
5. Two solutions have been titrated, one is light pink and one is a deep pink, which one do you think is more accurate? Why? (Think of what is happening in the reaction)
6. If it takes 30.05 mL of NaOH to react with 1.00 g of potassium hydrogen phthalate (KHP;  $\text{KHC}_8\text{H}_4\text{O}_4$ ), what is the concentration of NaOH?
7. How many trials must be run to calculate the average?
8. How reliable is a buret? How many significant figures result from a buret reading?
9. If a buret has an accuracy of  $\pm 0.2\%$ , what is the error range for a measurement of 1.56 mL?

### Pre-Lab Notebook Preparation:

Fill out all of the header information you will need.

Write out a brief statement of purpose

Write down any pertinent equations (think of the reaction between the unknowns and the base)