Chemistry 122 | Winter 2008 | Oregon State University
Exam 2 | February 28, 2008 | Dr. Richard Nafshun

Instructions: You should have with you several number two pencils, an eraser, your 3" x 5" note card, a calculator, and your University ID Card. If you have notes with you, place them in a sealed backpack and place the backpack OUT OF SIGHT or place the notes directly on the table at the front of the room.

Fill in the front page of the Scantron answer sheet with your last name, first name, middle initial, and student identification number. **Leave the class section number and the test form number blank.**

This exam consists of 25 multiple-choice questions. Each question has four points associated with it. Select the best multiple-choice answer by filling in the corresponding circle on the rear page of the answer sheet. If you have any questions before the exam, please ask. If you have any questions during the exam, please ask the proctor. Open and start this exam when instructed. When finished, place your Scantron form in the appropriate stack and present your University ID Card to the proctor. You may keep the exam packet, so please show your work and mark the answers you selected on it.

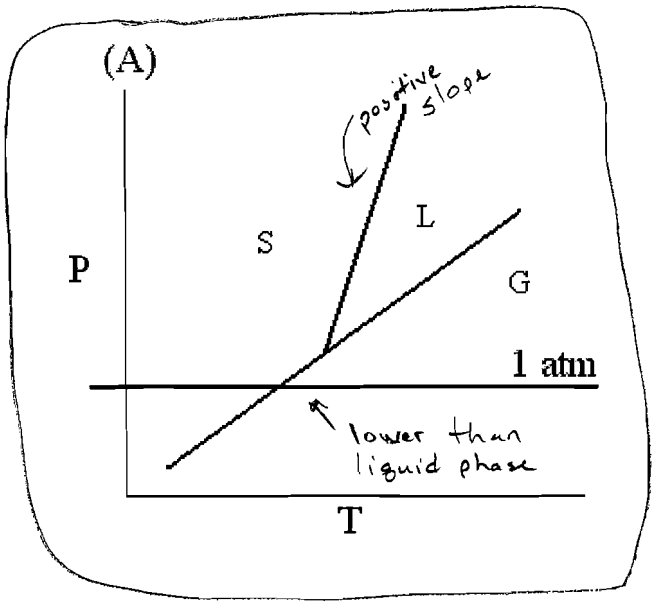
$R = 0.0821$ L•atm/mol•K | 760 mm Hg = 760 torr = 1 atm | m = mol/kg
M = mol/L | $\Delta T_f = imk_f$ | $\Delta T_b = imk_b$
IIV = nRT | $k_f(H_2O) = 1.86$ °C/m | $k_b(H_2O) = 0.512$ °C/m

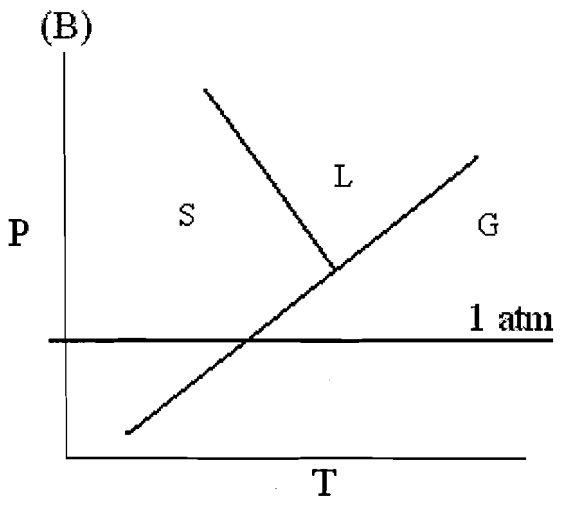
$\ln[\frac{A}{A_o}] = -kt$ | SC: $2r = s$ | BCC: $4r = s\sqrt{3}$ | FCC: $4r = s\sqrt{2}$

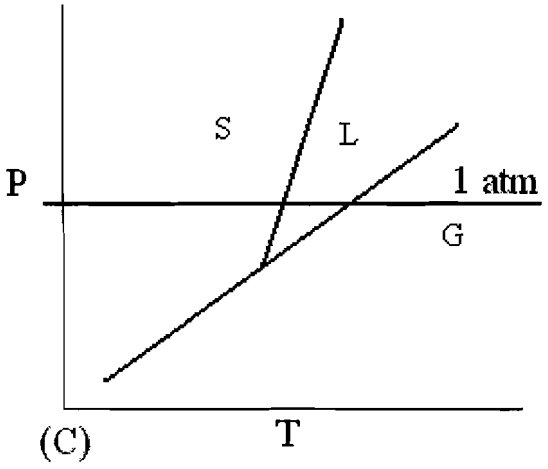
IA												IIIA	IVA	VA	VIA	VIIA	VIIIA
1 H Hydrogen 1.0079																	2 He Helium 4.0026
3 Li Lithium 6.941	4 Be Beryllium 9.01218											5 B Boron 10.81	6 C Carbon 12.011	7 N Nitrogen 14.0067	8 O Oxygen 15.9994	9 F Fluorine 18.9984	10 Ne Neon 20.179
11 Na Sodium 22.98977	12 Mg Magnesium 24.305	IIIB	IVB	VB	VIB	VIIB	VII			1B	IIB	13 Al Aluminum 26.9815	14 Si Silicon 28.0855	15 P Phosphorus 30.97376	16 S Sulfur 32.06	17 Cl Chlorine 35.453	18 Ar Argon 39.948
19 K Potassium 39.0983	20 Ca Calcium 40.08	21 Sc Scandium 44.9559	22 Ti Titanium 47.88	23 V Vanadium 50.9415	24 Cr Chromium 51.996	25 Mn Manganese 54.9380	26 Fe Iron 55.847	27 Co Cobalt 58.9332	28 Ni Nickel 58.70	29 Cu Copper 63.546	30 Zn Zinc 65.38	31 Ga Gallium 69.72	32 Ge Germanium 72.59	33 As Arsenic 74.9216	34 Se Selenium 78.96	35 Br Bromine 79.904	36 Kr Krypton 83.80
37 Rb Rubidium 85.4678	38 Sr Strontium 87.62	39 Y Yttrium 88.9059	40 Zr Zirconium 91.22	41 Nb Niobium 92.9064	42 Mo Molybdenum 95.94	43 Tc Technetium 98.906	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.9055	46 Pd Palladium 106.4	47 Ag Silver 107.868	48 Cd Cadmium 112.41	49 In Indium 114.82	50 Sn Tin 118.69	51 Sb Antimony 121.75	52 Te Tellurium 127.60	53 I Iodine 126.9045	54 Xe Xenon 131.30
55 Cs Cesium 132.9054	56 Ba Barium 137.33	57–71 *Rare earths	72 Hf Hafnium 178.49	73 Ta Tantalum 180.9479	74 W Tungsten 183.85	75 Re Rhenium 186.207	76 Os Osmium 190.2	77 Ir Iridium 192.22	78 Pt Platinum 195.09	79 Au Gold 196.9665	80 Hg Mercury 200.59	81 Tl Thallium 204.37	82 Pb Lead 207.2	83 Bi Bismuth 208.9804	84 Po Polonium (209)	85 At Astatine (210)	86 Rn Radon (222)
87 Fr Francium (223)	88 Ra Radium 226.0254	89–103 †Actinides	104 Rf Rutherfordium (261)	105 Ha Hahnium (262)	106 Sg Seaborgium (263)	107 Ns Neilsbohrium (262)	108 Hs Hassium (265)	109 Mt Meitnerium (266)	110 ‡ (269)	111 ‡			114				→ Stable region?

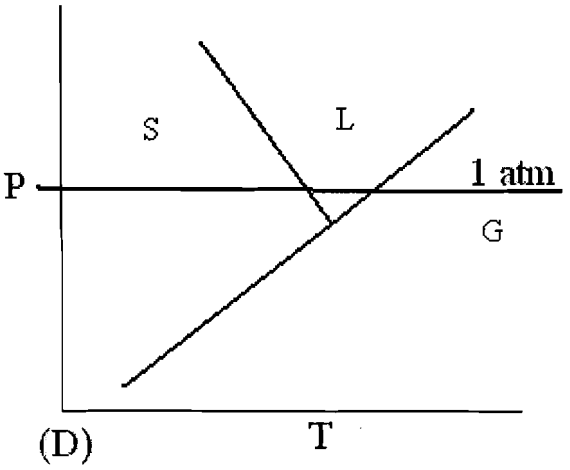
57 La Lanthanium 138.9055	58 Ce Cerium 140.12	59 Pr Praseodymium 140.9077	60 Nd Neodymium 144.24	61 Pm Promethium 145	62 Sm Samarium 150.4	63 Eu Europium 151.96	64 Gd Gadolinium 157.25	65 Tb Terbium 158.9254	66 Dy Dysprosium 162.50	67 Ho Holmium 164.9304	68 Er Erbium 167.26	69 Tm Thulium 168.9342	70 Yb Yterbium 173.04	71 Lu Lutetium 174.967
89 Ac Actinium 227.0278	90 Th Thorium 232.0381	91 Pa Protactinium 231.0359	92 U Uranium 238.029	93 Np Neptunium 237.0482	94 Pu Plutonium (244)	95 Am Americium (243)	96 Cm Curium (247)	97 Bk Berkelium (247)	98 Cf Californium (251)	99 Es Einsteinium (254)	100 Fm Fermium (257)	101 Md Mendelevium (258)	102 No Nobelium 259	103 Lr Lawrencium 262

1. Consider the phase diagrams below. Which diagram could correctly describe CO_2?







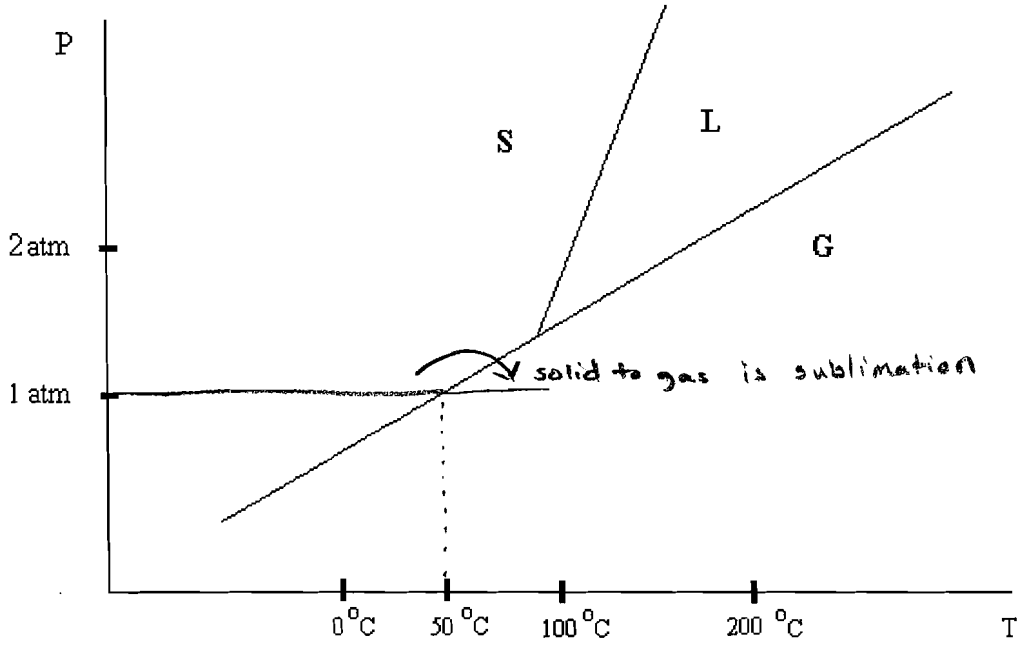


2. Which of the following is **false**?

(A) Graphite is a network covalent compound. True
(B) Diamond is a network covalent compound. True
(C) Network covalent compounds typically melt at higher temperatures than molecules. True
(D) Potassium chloride is a network covalent compound. ← KCl is an ionic compound
(E) Carbon dioxide is a molecule. True

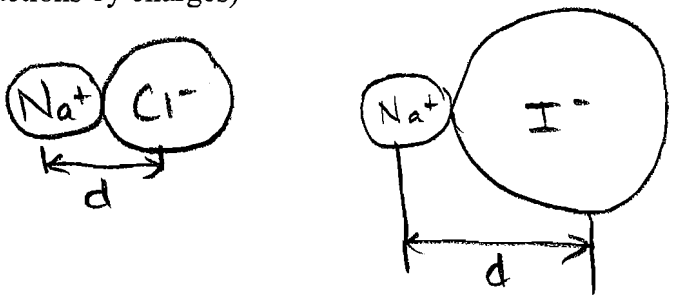
3. Consider the phase diagram below for compound foolsgoldium. The temperature of sublimation at 1 atm is:

(A) 0 °C
(B) 50 °C
(C) 100 °C
(D) 200 °C
(E) 212 °C



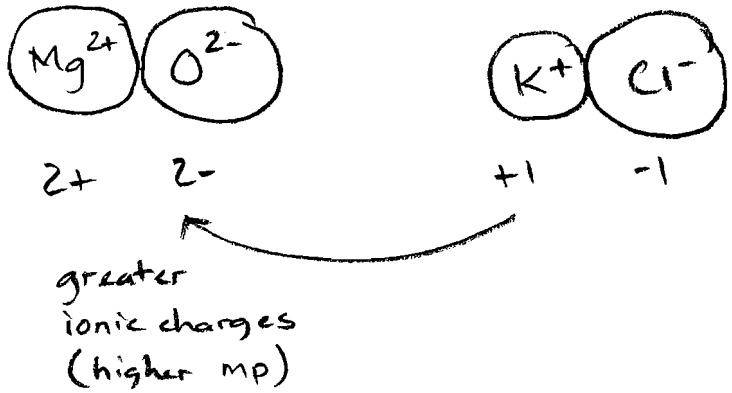
4. Sodium chloride melts at 804 °C. Sodium iodide melts at 651 °C. The difference in melting points can be attributed to:

(A) Different intermolecular forces (dispersion, dipole-dipole, hydrogen bonding)
(B) Different ionic charges (+1, +2, +3, -1, -2, -3…)
(C) Different distances between nuclei (ionic size)
(D) Network covalent compounds
(E) One is a molecule (attractions by intermolecular forces), one is an ionic compound (attractions by charges)

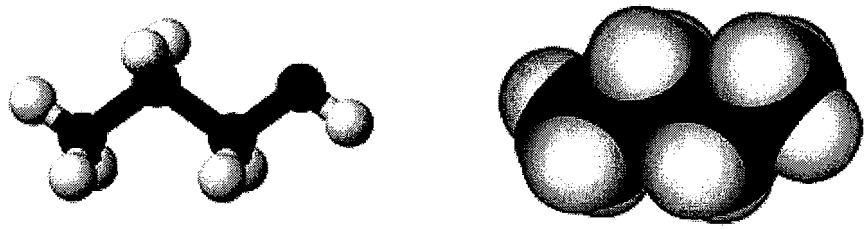


5. MgO melts near 2800 °C. KCl melts near 776 °C. The difference in melting points can be attributed to:

(A) Different intermolecular forces (dispersion, dipole-dipole, hydrogen bonding)
(B) Different ionic charges (+1, +2, +3, -1, -2, -3…)
(C) Different distances between nuclei (ionic size)
(D) Network covalent compounds
(E) The sheet-like structure



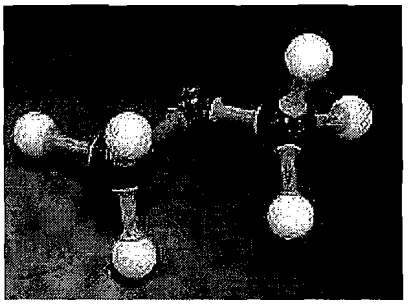
6. 1-propanol, $CH_3CH_2CH_2OH$, boils at +97.1 °C. Butane, $CH_3CH_2CH_2CH_3$ boils at −0.5 °C. The difference in melting points can be attributed to:



polar plus hydrogen bonding

(A) Different intermolecular forces (dispersion, dipole-dipole, hydrogen bonding)
(B) Different ionic charges (+1, +2, +3, -1, -2, -3…)
(C) Different distances between nuclei (ionic size)
(D) Network covalent compounds
(E) One is a molecule (attractions by intermolecular forces), one is an ionic compound (attractions by charges)

7. Consider dimethyl ether, CH_3OCH_3. The intermolecular forces present in CH_3OCH_3 are:



polar but no hydrogen bonding - all hydrogen atoms are bonded to carbon atoms, not to the oxygen atom

(A) Dispersion forces only.
(B) Dipole-dipole forces only.
(C) Dispersion forces and dipole-dipole forces.
(D) Dispersion forces, dipole-dipole forces, and hydrogen bonding.
(E) Hydrogen bonding only.

8. Consider He, CH_3OCH_3, CH_3CH_2OH, diamond, LiF, CsF, MgF_2, and AlN. Arranged in **increasing** melting point, these are:

	Lowest mp	**Highest mp**
(A)	He < CH_3OCH_3 < CH_3CH_2OH <LiF < CsF < MgF_2 < AlN < diamond	
(B)	He < CH_3OCH_3 < CH_3CH_2OH <CsF < LiF < MgF_2 < AlN < diamond	
(C)	He < CH_3CH_2OH < CH_3OCH_3 <CsF < LiF < MgF_2 < AlN < diamond	
(D)	He < CH_3CH_2OH < CH_3OCH_3 <LiF < CsF < MgF_2 < AlN < diamond	
(E)	He < CH_3OCH_3 < CH_3CH_2OH < AlN <LiF < CsF < MgF_2 < diamond	

Highest Melting Point →

Inert Gases	Polar Molecules	Molecules with hydrogen bonding	Ionic Solids	Network
He	CH_3OCH_3	CH_3CH_2OH	LiF, CsF, MgF_2, AlN	Diamond

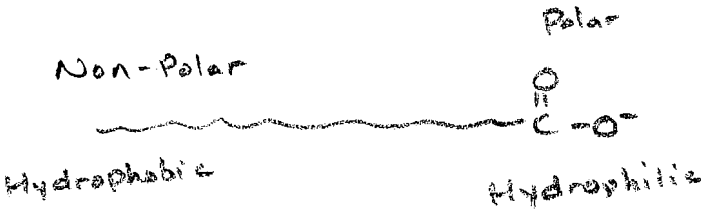
CsF LiF MgF_2 AlN
+1 -1 +1 -1 +2 -1 +3 -3

Li^+ F^- Cs^+ F^-

Both CsF and LiF are (+1) (-1)

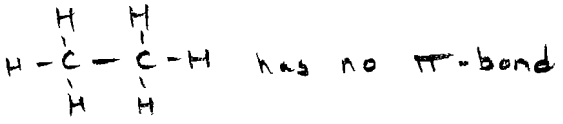
9. Which of the following has a hydrophilic end (polar, water-loving end) and a hydrophobic end (non-polar, water-fearing end) and has the ability to bridge water molecules to non-polar molecules?

(A) The polymer Teflon ($-CF_2CF_2-$)
(B) Graphite
(C) Ethane, CH_3CH_3
(D) Soap
(E) He

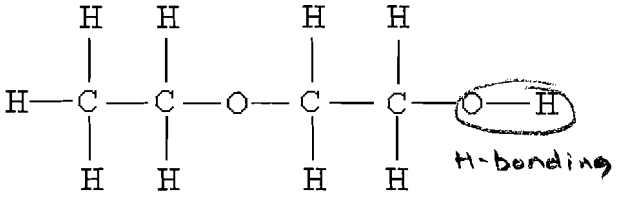


10. Which of the following cannot undergo free radical polymerization?

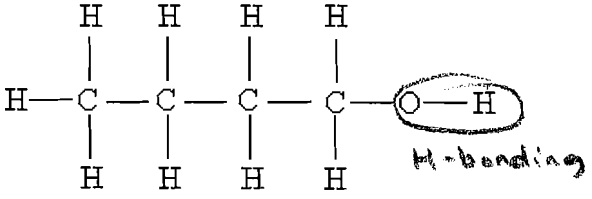
(A) C_2H_6
(B) C_2H_4
(C) C_2F_4
(D) C_2H_2



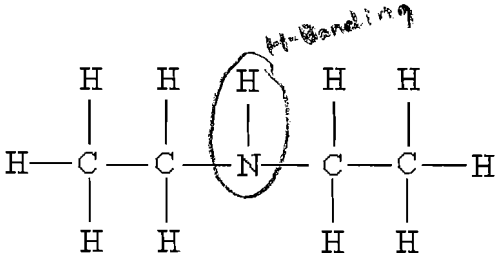
11. Which of the following molecules will **not** form hydrogen bonds?



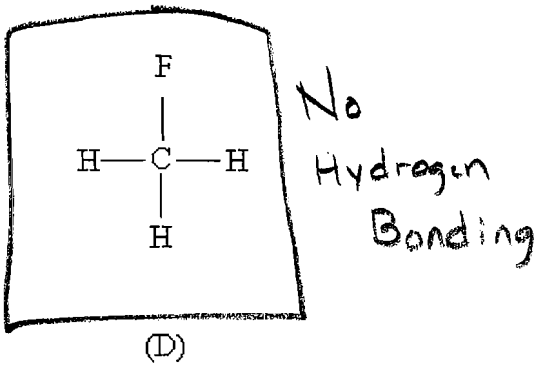
(A)



(B)



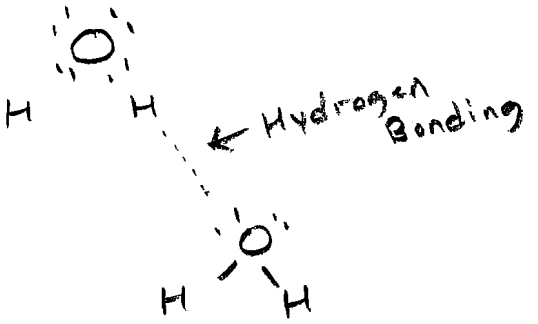
(C)



(D)

12. The intermolecular forces that are **most** significant in accounting for the high boiling point of liquid water relative to other substances of similar molecular weight is/are the:

(A) Dispersion forces
(B) Dipole-dipole interactions
(C) Hydrogen bonding
(D) Network covalent forces
(E) Ionic charges



13. Which has a higher melting point, sodium fluoride or aluminum fluoride? Why?

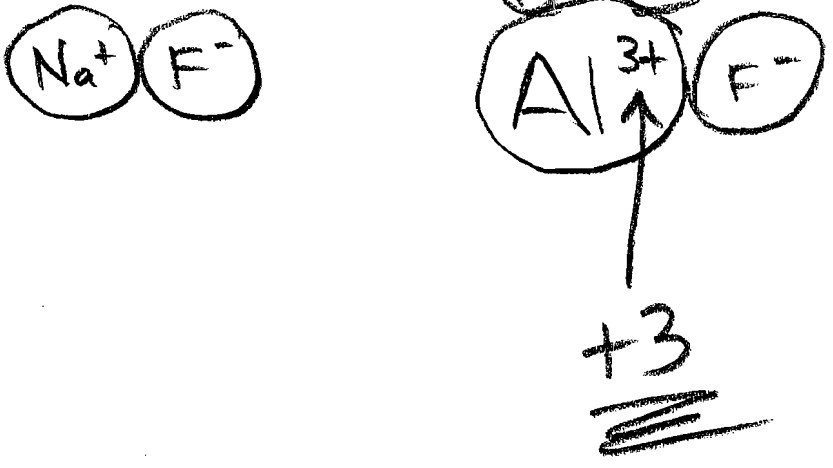
(A) sodium fluoride has a higher melting point because it has weaker dispersion forces than aluminum fluoride.

(B) sodium fluoride has a higher melting point because it has stronger dispersion forces than aluminum fluoride.

(C) aluminum fluoride has a higher melting point because it has stronger dispersion forces than sodium fluoride.

(D) aluminum fluoride has a higher melting point because it has a greater mass than sodium fluoride.

(E) aluminum fluoride has a higher melting point because it has greater ionic charges than sodium fluoride.

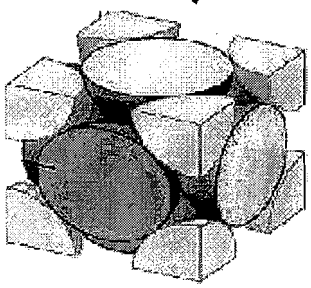


14. The equivalent number of atoms in the FCC unit cell is:

(A) 1
(B) 2
(C) 3
(D) 4
(E) 6

$8 \times \frac{1}{8} = 1$ (corners)
$6 \times \frac{1}{2} = 3$ (faces)
4

15. The structure below [from a *Course Worksheet*] represents:



(A) An SC unit cell
(B) A BCC unit cell
(C) A FCC unit cell
(D) A cell phone
(E) A prokaryotic cell

16. The cubic form for the fictitious element Nathanium (named for a CH 122 TA) is FCC. The atomic radius is 132.0 pm and the molar mass is 267.4 g/mol. The density of Nathanium is:
[1 m = 1 x 10^{12} pm 1 m = 100 cm]

(A) 34.2 g/cm^3
(B) 136.8 g/cm^3
(C) 8.55 g/cm^3
(D) 49.4 g/cm^3
(E) 2.03 g/cm^3

$d = \frac{m}{V}$

① Mass ⇒

$$267.4\ {}^{g}/_{mol}\left(\frac{1\ mol}{6.022\times10^{23}\ atoms}\right)\left(\frac{4\ atoms}{1\ FCC}\right) = 1.776\times10^{-21}\ g$$

② Volume ⇒ $V = s^3$

$r = 132.0\ pm$

$$132.0\ pm\left(\frac{1\ m}{1\times10^{12}\ pm}\right)\left(\frac{100\ cm}{1\ m}\right) = 1.32\times10^{-8}\ cm$$

$$s = \frac{4r}{\sqrt{2}} = \frac{(4)(1.32\times10^{-8}\ cm)}{\sqrt{2}}$$

$$s = 3.73\times10^{-8}\ cm$$

$$V = s^3 = (3.73\times10^{-8}\ cm)^3 = 5.20\times10^{-23}\ cm^3$$

$$d = \frac{m}{V} = \frac{1.776\times10^{-21}\ g}{5.20\times10^{-23}\ cm^3} = 34.2\ \frac{g}{cm^3}$$

Cl^- Cl^- Al^{3+} Cl^-

← i = 4 4 ions →

17. The freezing point of 0.500 m aqueous $AlCl_3$ is:

(A) -2.79 °C
(B) +2.79 °C
(C) +5.58 °C
(D) -5.58 °C
(E) -3.72 °C

$\Delta T_f = i m k_f = (4)(0.500\,m)\left(1.86 \frac{°C}{m}\right) = 3.72°C$

$T_f = 0°C - 3.72°C = -3.72°C$

18. Which of the following sets of compounds are expected to be soluble in water?

polar and ionic

(A) CH_4, CO_2, CF_4 ✗
(B) NaCl, CH_4, CH_3OCH_3 ✗
(C) NaCl, CH_3CH_2OH, NH_3
(D) NaCl, CCl_4, C_4H_{10} ✗

19. A student dissolves 12.000 g of an unknown polymer in 800 mL of water at 320 K. She measures the osmotic pressure to be 0.0677 mm Hg. What is the molar mass of the polymer?

(A) 2.71×10^6 g/mol
(B) 4.42×10^6 g/mol
(C) 1.73×10^5 g/mol
(D) 1.73×10^6 g/mol
(E) 2.26×10^6 g/mol

$\pi V = nRT$

$n = \frac{\pi V}{RT} = \frac{(0.0677 \text{ mm Hg} / 760 \text{ mm Hg} / 1 \text{ atm})(0.800 \text{ L})}{(0.08216 \frac{\text{L atm}}{\text{mol K}})(320 \text{ K})}$

$n = 2.71 \times 10^{-6} \text{ mol}$

$\text{Molar Mass} = \frac{g}{mol} = \frac{12.000 \text{ g}}{2.71 \times 10^{-6} \text{ mol}} = 4.42 \times 10^6 \frac{g}{mol}$

20. Generally, which of the following generally **does not** increase the reaction rate?

(A) Increase the temperature True
(B) Add a catalyst True
(C) Increase the activation energy False
(D) Increase the reactant concentrations True

$k = A e^{-\frac{E_a}{RT}}$

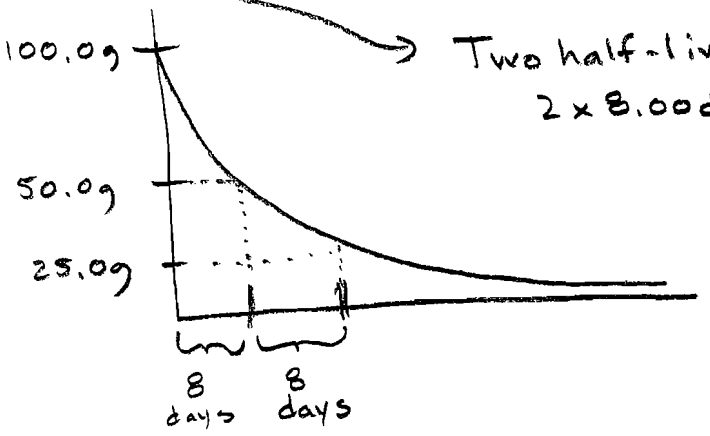
→ Decreasing E_a (with a catalyst) will increase the rate

21. The half-life is:

(A) The amount of time required for half the sample to decay
(B) 0.500 years
(C) The amount of time required for the entire sample to decay
(D) 42 years
(E) $\left(\frac{A}{A_o}\right)$

22. A student ( ) obtains a 100.0 gram sample of ^{131}I ($t_{1/2}$ = 8.00 days). How many grams of ^{131}I will remain after 16.00 days?

(A) 8.0 grams
(B) 16.0 grams
(C) 25.0 grams
(D) 50.0 grams
(E) 75.0 grams



23. A student ( ) obtains a 100.0 gram sample of ^{131}I ($t_{1/2}$ = 8.00 days). How long will it take so that only 10.0 grams of ^{131}I remain?

(A) 8.2 days
(B) 16.4 days
(C) 25.0 days
(D) 26.6 days
(E) 50.0 days

① Calc k

$$\ln\left[\frac{1}{2}\right] = -kt_{1/2}$$

$$\ln\left[\frac{1}{2}\right] = -k(8.00\,d)$$

$$k = 0.0866\,\frac{1}{d}$$

② Calc t

$$\ln\left[\frac{A}{A_o}\right] = -kt$$

$$\ln\left[\frac{10.0\,g}{100.0\,g}\right] = -\left(0.0866\,\frac{1}{d}\right)(t)$$

$$t = 26.6\,d$$

24. The following are initial rate data for: **A + B → C**

Experiment	Initial [A]	Initial [B]	Initial Rate
1	0.10	0.10	5.1
2	0.20	0.10	5.1
3	0.10	0.20	20.4

(A) The rate law is Rate = $k[A]^1[B]^2$.
(B) The rate law is Rate = $k[A]^0[B]^2$.
(C) The rate law is Rate = $k[A]^2[B]^0$.
(D) The rate law is Rate = $k[A]^2[B]^1$.
(E) The rate law is Rate = $k[A]^1[B]^4$.

Double [B] and the rate quadruples
$[B]^2$

Double [A] and no change in the rate
$[A]^0$

25. The Chemistry 122 final exam is Monday, March 17, 2008 at 4:00pm; yes, this is better than early in the morning. After the chemistry final I will be...

(A) Watching the C-SPAN baseball steroid hearings

(B) Making a poster for my dorm room wall that has the daily countdown to the new Indiana Jones movie (In theaters May 22)

(C) Two words: Tom and Gisele

(D) Preparing for my other seven final exams

(E) Re-watching The Oscars

[Any response will receive full credit; even no response]